Practice Problems: Atomic Mass Calculations II

See the complete solutions to these problems at https://www.youtube.com/watch?v=4bZhUEyVg0U.

1. Magnesium has three naturally occurring isotopes, 24 Mg (78.99%, 23.9850 amu), 25 Mg (10.00%, 24.9858 amu), and 26 Mg. Determine the percent abundance and isotopic mass of 26 Mg.

2. Boron has two naturally occurring isotopes. Find the percent abundances of ^{10}B and ^{11}B given the isotopic mass of $^{10}B = 10.0129$ amu and the isotopic mass of $^{11}B = 11.0093$ amu.



3.	Chlorine has two naturally occurring isotopes, ³⁵ Cl (34.9689 amu) and ³⁷ Cl (36.9659 amu). If chlorine has an average atomic mass of 35.4527 amu, what is the percent abundance of each isotope?